# **ATOMIC STRUCTURE**

- 1. The Nuclear Atom
- 2. Electron Orbits
- 3. Atomic Spectra
- 4. The Bohr Atom
- 5. Energy Level and Spectra
- 6. Correspondence Principle
- 7. Nuclear Motion
- 8. Atomic Excitation
- 9. The Laser

#### How atoms are excited?

- 1. Collisions.
- 2. Photon absorption.







#### How atoms are excited?

- 1. Collisions.
  - Part of their joint KE is absorbed by the atom.
  - The excited atom return to its ground by emitting one or more photons after on average 10<sup>-8</sup> s.
  - Discharge in rarefied gas occurs when an electric field accelerates electrons and atomic ions until their KE are sufficient to excite atoms they collide with.



### How atoms are excited?

### 2. Photon absorption.

- Atoms absorb a photon of light whose energy is the right amount to raise the atom to a higher energy level.
- This process explains the origin of absorption spectra.



### Franck-Hertz Experiment...

- Its an experiment to investigate energy levels inside atoms.
- It demonstrated the existence of atomic energy levels.
- The plate current was measured as a function of the accelerating potential.
- Two types of collisions occurred:
  - Elastic
  - Inelastic
- Critical potentials are due to atomic energy levels.



#### **Remember...** How atoms absorb and emit energy.

# A laser is a device that produces a light beam with some remarkable properties...

- 1. Very nearly monochromatic.
- 2. Coherent (all wave are in phase).
- 3. Collimated (hardly diverges).
- 4. Very intense.

Light Amplification by Stimulated Emission of Radiation

The key to the laser is the presence of a **metastable** state.



What are the kinds of transitions involving EM radiation that are possible between two energy levels?

- 1. Stimulated absorption.
- 2. Spontaneous emission.
- 3. Stimulated emission.



#### Three & four level laser...



### **Population inversion and optical pumping.**

#### **Remember...** Producing light waves in step.